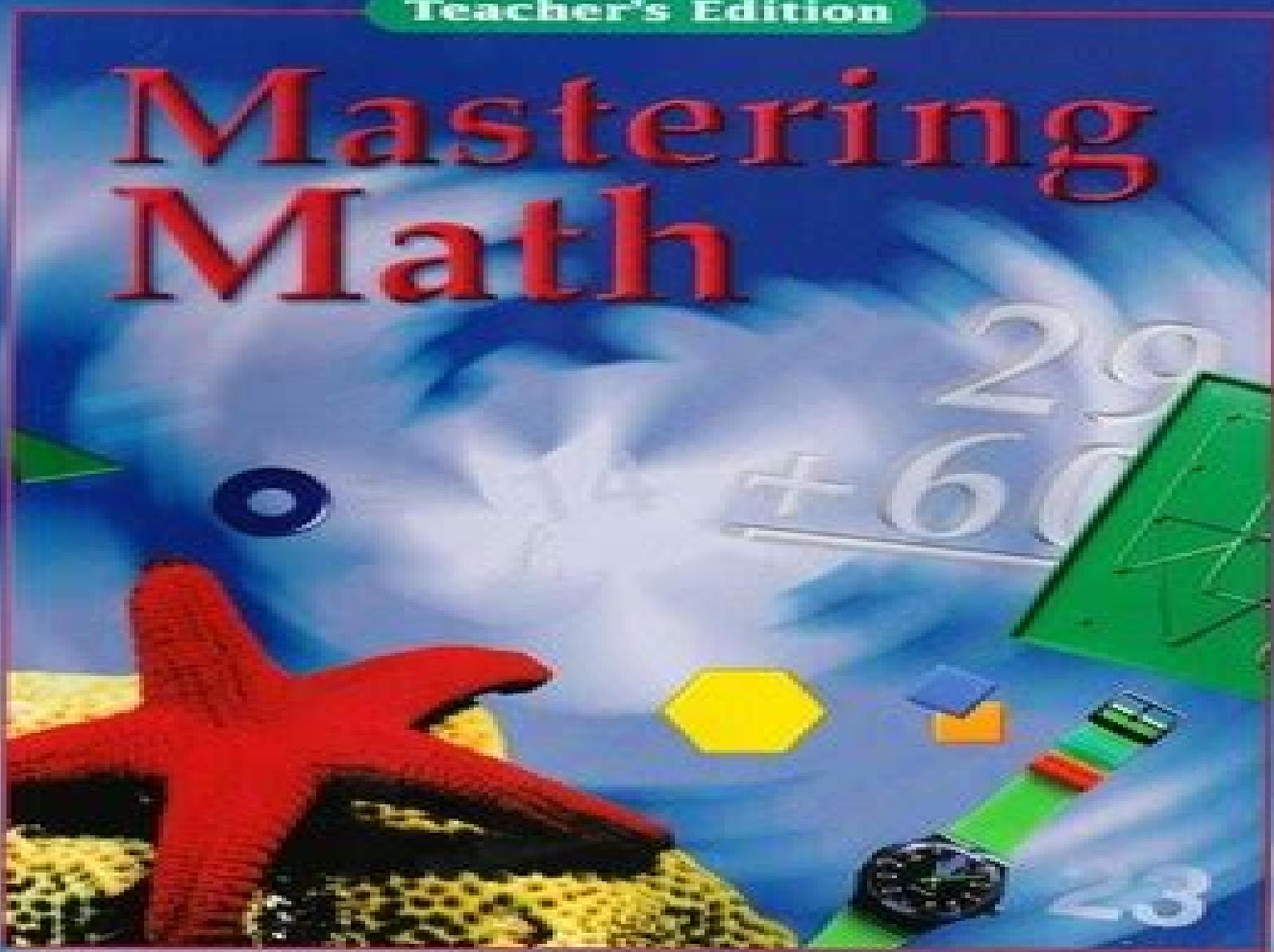


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web an ion that is found in both salts in a solution common ion effect the lowering of the solubility of an ionic compound as a result of the addition of a common ion is called this lower the smaller the numerical value of the solubility product constant the the solubility of the compound precipitate

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web hexane is a nonpolar liquid with a dipole moment of zero and therefore does not significantly interact with the ions of the nacl crystals 11 a fe no 3 3 is a strong electrolyte thus it should completely dissociate into fe 3 and no 3 ions therefore z best represents the solution b fe no 3 3 s fe 3 a q 3 no 3 a q

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web modern chemistry 149 chemical equilibrium chapter 18 review chemical equilibrium section 4 short answer answer the following questions in the space provided 1 match the solution type on the right to the corresponding relationship between the ion product and the  $K_{sp}$  for that solution listed on the left the ion product exceeds

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web here the value of concentration of water is taken as 0 36 instead of 0 16 divide the product of the concentration of water carbon monoxide and the equilibrium constant by the concentration of hydrogen instead of multiplying it correct d the value of  $K_{eq}$  is taken as 4 4 instead of 0 44