

## SOLUTIONS & SOLUBILITIES

### Solutions

A **solution** is a homogeneous mixture consisting of particles 0.1–2.0 nm in diameter. Homogeneous mixtures having larger particles (2–500 nm) are classified as **colloids**. *Suspensions* are mixtures with even larger particles, but they are not considered true solutions because they separate upon standing.

We usually think of solutions as solids dissolved in liquids, or perhaps a mixture of two liquids, but there are many other kinds of solutions as well. Some examples are provided in Table 11.1.

**TABLE 11.1** Some Different Kinds of Solutions

Kind of Solution	Example
Gas in gas	Air (O <sub>2</sub> , N <sub>2</sub> , Ar, and other gases)
Gas in liquid	Carbonated water (CO <sub>2</sub> in water)
Gas in solid	H <sub>2</sub> in palladium metal
Liquid in liquid	Gasoline (mixture of hydrocarbons)
Liquid in solid	Dental amalgam (mercury in silver)
Solid in liquid	Seawater (NaCl and other salts in water)
Solid in solid	Metal alloys, such as sterling silver (92.5% Ag, 7.5% Cu)

For solutions in which a gas or solid is dissolved in a liquid, the dissolved substance is called the **solute** and the liquid is called the **solvent**. When one liquid is dissolved in another, the minor component is usually considered the solute.

**Solution:** a homogeneous mixture containing particles the size of a typical ion or covalent molecule

**Colloid:** a homogeneous mixture containing particles with diameters in the range 2–500 nm

**Solute:** the dissolved substance in a solution

**Solvent:** the major component in a solution

# Solutions Solubilities 2vol Techniques

**Marcel A. Müller**



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