

# ALGEBRA I END OF COURSE STUDY GUIDE

## ALGEBRA I End of Course STUDY GUIDE

**Slope**  $y = mx + b$   
*m* is the slope  
*b* is the y-axis intercept

Find Slope given a table:



Find Slope given a graph:



Linear Functions:	Possible representations:
Save type gas (1 week)	100 miles
Silly string cost \$4.25 per can	1.25 per can
Charged \$4.25 every minute	4.25 per minute
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Represented by half price on a weekend	0.5 per hour
One person pays 1/2 admission, 2 = hours	0.5 per hour
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**Exponent Basics**  $x^a \cdot x^b = x^{a+b}$   $x^a \div x^b = x^{a-b}$   $x^0 = 1$

**Add Exponents & multiply numbers**

$$x^2 \cdot x^3 = x^5$$

$$x^3 \cdot x^2 = 2x \cdot 2x \cdot 2x \cdot x^2 = 8x^5$$

$$3x \cdot 2x^2 = 6x^3$$

$$x^2 \cdot x^3 = x^2 \cdot x^2 \cdot x^2 = x^6$$

$$(3 \text{ dots } 2)^2 = 3^2 \cdot 2^2 = 9 \cdot 4 = 36$$

$$4 \cdot x^2 \cdot y^3 = 4 \cdot x^2 \cdot y^3$$

$$(2x^2 \cdot 3x^2)^2 = (6x^4)^2 = 36x^8$$

**Subtract Exponents & Divide numbers**

$$\frac{x^5}{x^2} = x^{5-2} = x^3$$

$$\frac{27x^3y^3}{3x^2y^2} = \frac{9x^1y^1}{1} = 9xy$$

$$\frac{(-6x^2)(2x^3)}{-18x^5} = \frac{-12x^5}{-18x^5} = \frac{2}{3}$$

Independent	Dependent
$x$	$f(x)$
Domain	Range
$2x - 4$	$y =$
$8h$	$V =$
$ax^2$	$A =$
Number of items	Total cost \$
Number of hours worked	Amount of money \$ earned

**Solve for 'y'**

$$3x + 4y = 24$$

$$-3x$$

$$4y = -3x + 24$$

$$\frac{4y}{4} = \frac{-3x + 24}{4}$$

$$y = -\frac{3}{4}x + 6$$

Steps:

- Add positive or negative to opposite side of  $y$ .
- Cancel out  $y$  then bring down carefully.
- Divide by the number to front of  $y$ .
- Rewrite.

## Distributive property

$$4(x + 3) = 4x + 12$$

$$-3x(7x - 4) + 6x = (-21x^2 + 12x) + 6x$$

$$= -21x^2 + 18x$$

$$-21x^2 + 12x + 6x = 18x - 21x^2$$

$$21x^2 + 18x - 21x^2$$

$$\frac{2}{5} (15a^2b - 40ab^2) - \frac{2}{3} (30ab^2 - 6a^2b)$$

$$6a^2b - 16ab^2 - 20ab^2 + 4a^2b$$

$$10a^2b - 32ab^2$$

Combine Like Terms  
 Same Variables with the Same Exponents

$$2x + 3x = 5x$$

$$3x + 2x = 5x$$

$$2x + 3x = 5x$$

$$3x + 2x = 5x$$



## Transformations of Quadratic Functions

**Vertical Shift**  
 Replacing  $x^2$  with  $x^2 + c$  results in shift of  $c$  units **up**.  
 Replacing  $x^2$  with  $x^2 - c$  results in shift of  $c$  units **down**.



**Horizontal Compression**  
 Changing  $x^2$  to  $(\frac{1}{2}x)^2$  compresses the parabola horizontally, making it **NARROWER**.

**Horizontal Stretch**  
 Changing  $x^2$  to  $(2x)^2$  stretches the parabola horizontally, making it **WIDER**.

# Study Guide For Intermediate Algebra

**Chao Zhang**



## **Study Guide For Intermediate Algebra:**

## **Study Guide For Intermediate Algebra** Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Study Guide For Intermediate Algebra**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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limiting reagent if 8.52 g are actually formed what is the  
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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $Al + O_2 \rightarrow Al_2O_3$  4 Al 3  $O_2 \rightarrow 2 Al_2O_3$  b determine the theoretical yield of  $Al_2O_3$  160.0 g Al  $\times 1$  mol Al  $\times 2$  mol  $Al_2O_3 \times 101.964$  g  $Al_2O_3 = 302.3$  g  $Al_2O_3$  26.982 g Al 4 mol Al 1 mol  $Al_2O_3$  c determine the percent yield yield actual  $\times 100 = 260.0$  g  $\times 100 = 86.01$  theoretical 302.3 g

**7 2 limiting reagent and reaction yields chem 1114** - Jul 02 2022

web 7 Na 2 C 2 O 4 is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H 2 are left over 13

6 5 limiting reagent and percent yield chemistry libretexts - Aug 15 2023

web learning objectives define and determine theoretical yields actual yields and percent yields identify a limiting reagent from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s remains when the reaction is complete

**limiting reactant and reaction yields article khan academy** - Jan 08 2023

web the percent yield is calculated as follows text percent yield  $\frac{\text{text actual yield}}{\text{text theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

**8 5 limiting reactant theoretical yield and percent yield** - Dec 07 2022

web may 20 2018 learning objectives identify the limiting reactant limiting reagent in a given chemical reaction calculate how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is complete

**12 3 limiting reagent and percent yield worksheet answer key** - Jan 28 2022

web limiting reagent and percent yield practice worksheet answer key limiting reactant yield practice worksheet 1 methanol  $CH_3OH$  can be produced by the following reaction  $2H_2 + CO \rightarrow CH_3OH$  calculate the theoretical yield of  $CH_3OH$  if 68.5 g of CO is

reacted with 8.6 g of  $H_2$

limiting reactant and percent yield worksheet with key - Oct 05 2022

web honors chemistry 1b limit reactant and percent yield worksheet with excess calculation name sto 4 solve stoichiometric problems from a balanced chemical equation 3 sto 5 differentiate between the actual yield and theoretical yield of a chemical reaction

*limiting reactant and percent yield practice hs science* - Jun 01 2022

web limiting reactant and percent yield practice name 1 consider the following reaction  $NH_4NO_3 + Na_3PO_4 \rightarrow NH_4_3PO_4 + NaNO_3$  which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

*limiting reagents and percentage yield worksheet answers* - Sep 04 2022

web limiting reagents and percentage yield worksheet 1 consider the reaction  $I_2O_5 + 5CO \rightarrow I_2 + 5CO_2$  a 80.0 grams of iodine v oxide  $I_2O_5$  reacts with 28.0 grams of carbon monoxide  $CO$  determine the mass of iodine  $I_2$  which could be produced  
80 g  $I_2O_5$  1 mol  $I_2O_5$  1 mol  $I_2$  xs 1 333.8 g  $I_2O_5$  1 mol  $I_2O_5$  28 g  $CO$  1 mol  $CO$

**7 3 limiting reactant and percent yield problems** - Jun 13 2023

web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3