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**Judd E. Hollander**



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which is equal to 22.4 l for 1 mole of any ideal gas at a temperature equal to 273.15 K and a pressure

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web the molar volume will be the volume occupied by one molar mass 32 g of the liquid expressing the density in liters instead of ml we have  $v = \frac{m}{\rho}$   $\frac{32 \text{ g}}{1.790 \text{ g l}} = 17.88 \text{ l mol}^{-1}$  nonnumber

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web zero degrees Celsius or 273 K is defined as standard temperature and one atmosphere or 760 mm Hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure STP is 22.4 l

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