

# THE WEAKER VESSEL



# Weaker Vessel

**DJ Losen**



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web limiting reactant theoretical and percentage yields key  $4 \text{KO}_2 + 2 \text{H}_2\text{O} + \text{I}_4\text{KOH} \rightarrow 3 \text{O}_2 + \text{G}$  if a reaction vessel contains 0

mol ko 2 and 0 mol h 2 o what is the limiting reactant

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$$2\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole h<sub>2</sub>s and multiply it with 2 mole h<sub>2</sub>o over 2 mole h<sub>2</sub>s from the equation 8 3 mole h<sub>2</sub>s x 2 mole h<sub>2</sub>o / 2 mole h<sub>2</sub>s x 18g h<sub>2</sub>o = 149 4g h<sub>2</sub>o is your theoretical yield so percent yield = 137 1g h<sub>2</sub>o actual yield / 149 4g h<sub>2</sub>o x 100 = 91 77 problems 6 8 are similar to problem 5

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80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  xs 1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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