

# Sync vs Async Communication

Request and wait for a response  
Fire and forget with messages/events

@lorenz/23

## Synchronous



Request-response model

Send a request and wait for some response  
Example: API requests

## Asynchronous



Fire and Forget model

Send message/event and forget  
Example: Event Driven Architecture

# Async Communication For Beginners

**Charles M. Tolbert**



## **Async Communication For Beginners:**

*Beginning JavaScript* Jeremy McPeak, 2015-02-18 The bestselling JavaScript guide updated with current features and best practices *Beginning JavaScript* 5th Edition shows you how to work effectively with JavaScript frameworks functions and modern browsers and teaches more effective coding practices using HTML5 This new edition has been extensively updated to reflect the way JavaScript is most commonly used today introducing you to the latest tools and techniques available to JavaScript developers Coverage includes modern coding practices using HTML5 markup the JSON data format DOM APIs the jQuery framework and more Exercises with solutions provide plenty of opportunity to practice and the companion website offers downloadable code for all examples given in the book Learn JavaScript using the most up to date coding style Understand JSON functions events and feature detection Utilize the new HTML5 elements and the related API Explore new features including geolocation local storage and more JavaScript has shaped the Web from a passive medium into one that is rich dynamic and interactive No matter the technology on the server side it s JavaScript that makes it come alive in the browser To learn JavaScript the way it s used today *Beginning JavaScript* 5th Edition is your concise guide

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web identifying chemical reaction types key part 1 identify the type of reaction 1  $\text{Na}_3\text{PO}_4 + 3\text{KOH} \rightarrow 3\text{NaOH} + \text{K}_3\text{PO}_4$  double displacement 2  $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2\text{LiCl}$  double displacement 3  $\text{C}_6\text{H}_{12} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$  combustion 4  $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$  single displacement 5  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$  decomposition 6  $\text{P}_4 + 3\text{O}_2 \rightarrow 2\text{P}_2\text{O}_3$

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web chemical equations can also be used to represent physical processes write a chemical reaction for the boiling of water including the proper phase labels chemical equations can also be used to represent physical processes write a chemical reaction for the freezing of water including the proper phase labels explain why  $4\text{Na} + \text{S} \rightarrow 2\text{Na}_2\text{S}$

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web name score date balancing and identifying chemical reactions a identify each type of chemical reactions 1  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$  2  $\text{Sr} + \text{AgNO}_3 \rightarrow \text{Sr}(\text{NO}_3)_2 + 2\text{Ag}$  3  $\text{LiBr} + \text{F}_2 \rightarrow \text{LiF} + \text{Br}_2$  4  $2\text{Na} + \text{F}_2 \rightarrow 2\text{NaF}$  5  $\text{F}_2 + \text{Br}_2 \rightarrow 2\text{BrF}$  b balance the chemical reactions and identify each type

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web 1  $3\text{NaBr} + \text{H}_3\text{PO}_4 \rightarrow \text{Na}_3\text{PO}_4 + 3\text{HBr}$  type of reaction double displacement 2  $3\text{Ca}(\text{OH})_2 + 2\text{Al}_2(\text{SO}_4)_3 \rightarrow 3\text{CaSO}_4 + 2\text{Al}(\text{OH})_3$  type of reaction double displacement 3  $3\text{Mg} + \text{Fe}_2\text{O}_3 \rightarrow 2\text{Fe} + 3\text{MgO}$  type of reaction single displacement 4  $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$  type of reaction combustion 5  $2\text{PbSO}_4 + 2\text{PbSO}_3 + \text{O}_2$  type of reaction

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web a chemical reaction is a process in which one or more substances are converted to one or more different substances the starting substances are called the reactants and the new substances that form are called the products

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web study guide chemical reactions 1 give an example of a chemical reaction leaves turning color fireworks food spoiling 2 differentiate between a physical and chemical change physical does not change the identity of the matter can be changed back easily chemical new substance is created with new physical and chemical