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web these mass relationships made through moles are called stoichiometry gk stoicheon element metry measure using mole and mass relationships we can calculate the mass of product that should be produced from a given amount of reactant when it is completely consumed in the reaction

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web sep 28 2022 i write down balanced molecular equations for the chemical changes ii write down the number of moles below the formula of each of the reactants and products iii write down the relative masses of the reactants and products with the help of formula below the respective formula these shall be the theoretical amounts of reactants and

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web recall that the molar mass can be determined from a chemical formula and used as a conversion factor we can add that conversion factor as another step in a calculation to make a mole mass calculation where we start with a given number of moles of a substance and calculate the mass of another substance involved in the chemical equation or vice

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web mass 32 g mass 16 g mass 3 2 g 2 for jars d e f use the number of molecules in the circle to determine the number of moles of o 2 in the jar then calculate the mass of the gas in the jar d e f 12 mol o 2 2 mol o 2 6 mol o 2 mass 384 g mass 64 g mass 192 g 3 jars g h and i contain ammonia gas nh 3

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web solution average atomic mass of li 6.941 amu x 6.0151 amu 1 x 7.0160 amu 6.941 1.0009x 7.0160 1.0009x 0.075 x 0.075 x 0.075 corresponds to a natural abundance of 6.0151 percent the natural abundance of 7li is x 0.925 or 92.5 percent 3.7 the unit factor required is 6.022

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web a molar mass li₂co₃ 2 6.941 g 12.01 g 3 16.00 g 73.89 g b molar mass cs₂ 12.01 g 2 32.07 g 76.15 g c molar mass chcl₃ 12.01 g 1 008 g 3 35.45 g 119.37 g d molar mass c₆h₈o₆ 6 12.01 g 8 1.008 g 6 16.00 g 176.12 g e molar mass kno₃ 39 10 g 14.01 g 3 16.00 g 101.11 g

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web it is a small step from mole mass calculations to mass mass calculations if we start with a known mass of one substance

in a chemical reaction instead of a known number of moles we can calculate the corresponding masses of other substances in the reaction

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