

Standard Reduction Potentials in Aqueous Solution at 25° C

| Half-reaction | | $E^\circ(\text{V})$ |
|---|--|---------------------|
| $\text{F}_2(\text{g}) + 2 e^-$ | $\rightarrow 2 \text{F}^-$ | 2.87 |
| $\text{Co}^{3+} + e^-$ | $\rightarrow \text{Co}^{2+}$ | 1.82 |
| $\text{Au}^{3+} + 3 e^-$ | $\rightarrow \text{Au}(\text{s})$ | 1.50 |
| $\text{Cl}_2(\text{g}) + 2 e^-$ | $\rightarrow 2 \text{Cl}^-$ | 1.36 |
| $\text{O}_2(\text{g}) + 4 \text{H}^+ + 4 e^-$ | $\rightarrow 2 \text{H}_2\text{O}(\text{l})$ | 1.23 |
| $\text{Br}_2(\text{l}) + 2 e^-$ | $\rightarrow 2 \text{Br}^-$ | 1.07 |
| $2 \text{Hg}^{2+} + 2 e^-$ | $\rightarrow \text{Hg}_2^{2+}$ | 0.92 |
| $\text{Hg}^{2+} + 2 e^-$ | $\rightarrow \text{Hg}(\text{l})$ | 0.85 |
| $\text{Ag}^+ + e^-$ | $\rightarrow \text{Ag}(\text{s})$ | 0.80 |
| $\text{Hg}_2^{2+} + 2 e^-$ | $\rightarrow 2 \text{Hg}(\text{l})$ | 0.79 |
| $\text{Fe}^{3+} + e^-$ | $\rightarrow \text{Fe}^{2+}$ | 0.77 |
| $\text{I}_2(\text{s}) + 2 e^-$ | $\rightarrow 2 \text{I}^-$ | 0.53 |
| $\text{Cu}^+ + e^-$ | $\rightarrow \text{Cu}(\text{s})$ | 0.52 |
| $\text{Cu}^{2+} + 2 e^-$ | $\rightarrow \text{Cu}(\text{s})$ | 0.34 |
| $\text{Cu}^{2+} + e^-$ | $\rightarrow \text{Cu}^+$ | 0.15 |
| $\text{Sn}^{4+} + 2 e^-$ | $\rightarrow \text{Sn}^{2+}$ | 0.15 |

Standard Potentials In Aqueous Solution

S Ben Porath



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