

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's $\delta+$ region is attracted to a $\delta-$ region in another molecule.
3. Form between water molecules (Where are the $\delta+$ and $\delta-$ regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

Study Guide For Changs Chemistry

Raymond Chang



Study Guide For Changs Chemistry:

Student Study Guide for Chang Chemistry With Advanced Topics Raymond Chang, 2015-08-03 *Student Study Guide for Chemistry* Kenneth Goldsby, Raymond Chang, 2012-01-18 This valuable ancillary contains material to help the student practice problem solving skills For each section of a chapter the author provides study objectives and a summary of the corresponding text Following the summary are sample problems with detailed solutions Each chapter has true false questions and self test with all answers provided at the end of the chapter A Study Guide to Accompany Raymond Chang, Chemistry, Second Edition Kenneth Watkins, Raymond Chang, 1981 Student Study Guide to Accompany Chemistry Kenneth W. Watkins, Raymond Chang, 2001-02 by Kenneth W Watkins Colorado State University This valuable ancillary contains material to help the student organize study time practice problem solving skills and complete self tests For each section of a chapter the author provides study objectives and a summary of the corresponding text Following the summary are sample problems with detailed solutions Each chapter has true false questions and a self test with all answers provided at the end of the chapter *Study Guide for Chang's Chemistry* Kenneth W. Watkins, 1988 **Student Study Guide to accompany Chemistry** Raymond Chang, 2009-01-15 This valuable ancillary contains material to help the student practice problem solving skills For each section of a chapter the author provides study objectives and a summary of the corresponding text Following the summary are sample problems with detailed solutions Each chapter has true false questions and self test with all answers provided at the end of the chapter **Chang, AP Focus Review Guide** Raymond Chang, 2015-01-13

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web the relative electronegativities of the bonded atoms is h c cl and so the bond moments all point toward the cl end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as BF_3 trigonal planar CH_4 tetrahedral PF_5 trigonal bipyramidal and SF_6

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