

Physical Chemistry Formula Revision

Basic Concepts

$$\text{Average atomic mass} = \frac{a_1x_1 + a_2x_2 + a_3x_3 + \dots}{100}$$

a_1, a_2, a_3, \dots - atomic mass of isotopes

x_1, x_2, x_3, \dots - percentage abundance

Mole concept

$$\text{no. of moles, } n = \frac{\text{given mass}}{\text{molar mass}} = \frac{W}{M}$$

$$n = \frac{\text{no. of particles}}{\text{Avogadro's No.}} = \frac{N}{6.022 \times 10^{23}}$$

$$n = \frac{\text{vol. of gas}}{\text{vol. of gas at STP}} = \frac{V}{22.4 \text{ L}}$$

Concentration Terms

$$\text{Percentage, } \% = \frac{W_2 \text{ or } V_2}{W_1 \text{ or } V_1} \times 100$$

$$\text{Molarity, } M = \frac{n_2}{V(l)} = \frac{W_2 \times 1000}{M_2 \times V(\text{ml})}$$

$$\text{Normality, } N = \frac{n_2 \times 1000}{V(l)} = \frac{W_2 \times 1000}{E_2 \times V(\text{ml})}$$

$\hookrightarrow \frac{M_2}{n_f}$

$$\text{parts per million, ppm} = \frac{W_2}{W_1} \times 10^6$$

$$\text{Molality, } m = \frac{n_2}{M_2(\text{kg})} = \frac{W_2 \times 1000}{M_2 \times W_1(\text{g})}$$

$$\text{Mole fraction, } X_2 = \frac{n_2}{n_1 + n_2 + n_3 + \dots}$$

Stoichiometry



$$\frac{\text{moles of A}}{a} = \frac{\text{moles of B}}{b} = \frac{\text{moles of C}}{c}$$

Limiting R_2 : given amt. \rightarrow lowest stoichiometric coeff.

Atomic Structure

Planck's Quantum Theory

$$E = \frac{hc}{\lambda} \quad E(\text{eV}) = \frac{12400}{\lambda(\text{\AA})}$$

Photo Electric Effect

$$K_{\text{max}} = h\nu - \phi + K_{\text{max}}$$

$$e = h\nu - \frac{m v_{\text{max}}^2}{2}$$



De Broglie's Hypothesis

$$m v = \frac{h}{\lambda} \rightarrow \lambda = \frac{h}{m v} \quad \lambda = \frac{h}{\sqrt{2mE}}$$

Bohr's Atomic model

$$m v r = \frac{n h}{2\pi} \quad r = 0.53 \times \frac{n^2}{Z} \text{ \AA}$$

$$r = 2.18 \times 10^{-8} \frac{n^2}{Z} \text{ m}$$

$$E_n = -13.6 \times \frac{Z^2}{n^2} \text{ eV} \quad T.E. = E_2 - E_1 = -0.68 \text{ eV}$$

$$\frac{1}{\lambda} = R_H Z^2 \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right] \quad R_H = 1.097 \times 10^7 \text{ m}^{-1}$$

max no. of spectral lines when e^- returns from $n_2 = n_1 + 1$ to $n_1 = n_1(n_1 - 1) \frac{2}{2}$

Heisenberg's uncertainty principle: $\Delta x \Delta p \geq \frac{h}{4\pi}$

Quantum Model

$$\text{total nodes} = n - 1$$

$$\text{angular nodes} = l$$

$$\text{radial nodes} = n - l - 1$$

Angular momentum

$$\text{orbital} = \sqrt{l(l+1)} \frac{h}{2\pi}$$

$$\text{spin} = \sqrt{s(s+1)} \frac{h}{2\pi}$$

Structure Of Physical Chemistry

Peter William Atkins, Julio De Paula



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good day and the water in order but experience has taught most of us that the good days are in the minority and that as is the case with our rapid running streams such as many of our northern streams are the water is either too large or too small unless as previously remarked you live near at hand and can catch it at its best A common belief in regard to loch fishing is that the tyro and the experienced angler have nearly the same chance in fishing the one from the stern and the other from the bow of the same boat Of all the absurd beliefs as to loch fishing this is one of the most absurd Try it Give the tyro either end of the boat he likes give him a cast of ally flies he may fancy or even a cast similar to those which a crack may be using and if he catches one for every three the other has he may consider himself very lucky Of course there are lochs where the fish are not abundant and a beginner may come across as many as an older fisher but we speak of lochs where there are fish to be caught and where each has a fair chance Again it is said that the boatman has as much to do with catching trout in a loch as the angler Well we dont deny that In an untried loch it is necessary to have the guidance of a good boatman but the same argument holds good as to stream fishing

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