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Two Days To Remember

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 $3 \text{ hbr} + 1 \text{ khco}_3 \rightarrow 3 \text{ h}_2\text{o} + 1 \text{ kbr} + 1 \text{ co}_2$
 $3 \text{ ga} + 3 \text{ na}_2\text{so}_3 \rightarrow 3 \text{ ga}_2 + 3 \text{ so}_3 + 6 \text{ nabr}$
 $3 \text{ sno} + 2 \text{ nf}_3 \rightarrow 3 \text{ snf}_2 + 1 \text{ n}_2\text{o}_3$
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 $3 \text{ ga} + 3 \text{ na}_2\text{so}_3 \rightarrow 3 \text{ ga}_2 + 3 \text{ so}_3 + 6 \text{ nabr}$
 $3 \text{ sno} + 2 \text{ nf}_3 \rightarrow 3 \text{ snf}_2 + 1 \text{ n}_2\text{o}_3$
 using the following equation
 $2 \text{ naoh} + \text{ h}_2\text{so}_4 \rightarrow 2 \text{ h}_2\text{o} + \text{ na}_2\text{so}_4$

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 $2 \text{ zncl}_x + 2 \text{ cu} \rightarrow$
 how many moles of $\text{ce} + \text{zncl}_2 + \text{zncl}_x + 2$ will be produced from 23.0 g of $\text{ce} + \text{zn}$ assuming $\text{ce} + \text{cucl}_2 + \text{cucl}_x + 2$ is available in excess

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 $3 \text{ hbr} + 1 \text{ khco}_3 \rightarrow 3 \text{ h}_2\text{o} + 1 \text{ kbr} + 1 \text{ co}_2$
 $3 \text{ ga} + 3 \text{ na}_2\text{so}_3 \rightarrow 3 \text{ ga}_2 + 3 \text{ so}_3 + 6 \text{ nabr}$
 $3 \text{ sno} + 2 \text{ nf}_3 \rightarrow 3 \text{ snf}_2 + 1 \text{ n}_2\text{o}_3$
 using the following equation
 $2 \text{ naoh} + \text{ h}_2\text{so}_4 \rightarrow 2 \text{ h}_2\text{o} + \text{ na}_2\text{so}_4$

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