

# Standard Reduction Potentials in Aqueous Solution at 25° C

Half-reaction		$E^\circ(\text{V})$
$\text{F}_2(\text{g}) + 2 e^-$	$\rightarrow 2 \text{F}^-$	2.87
$\text{Co}^{3+} + e^-$	$\rightarrow \text{Co}^{2+}$	1.82
$\text{Au}^{3+} + 3 e^-$	$\rightarrow \text{Au}(\text{s})$	1.50
$\text{Cl}_2(\text{g}) + 2 e^-$	$\rightarrow 2 \text{Cl}^-$	1.36
$\text{O}_2(\text{g}) + 4 \text{H}^+ + 4 e^-$	$\rightarrow 2 \text{H}_2\text{O}(\text{l})$	1.23
$\text{Br}_2(\text{l}) + 2 e^-$	$\rightarrow 2 \text{Br}^-$	1.07
$2 \text{Hg}^{2+} + 2 e^-$	$\rightarrow \text{Hg}_2^{2+}$	0.92
$\text{Hg}^{2+} + 2 e^-$	$\rightarrow \text{Hg}(\text{l})$	0.85
$\text{Ag}^+ + e^-$	$\rightarrow \text{Ag}(\text{s})$	0.80
$\text{Hg}_2^{2+} + 2 e^-$	$\rightarrow 2 \text{Hg}(\text{l})$	0.79
$\text{Fe}^{3+} + e^-$	$\rightarrow \text{Fe}^{2+}$	0.77
$\text{I}_2(\text{s}) + 2 e^-$	$\rightarrow 2 \text{I}^-$	0.53
$\text{Cu}^+ + e^-$	$\rightarrow \text{Cu}(\text{s})$	0.52
$\text{Cu}^{2+} + 2 e^-$	$\rightarrow \text{Cu}(\text{s})$	0.34
$\text{Cu}^{2+} + e^-$	$\rightarrow \text{Cu}^+$	0.15
$\text{Sn}^{4+} + 2 e^-$	$\rightarrow \text{Sn}^{2+}$	0.15

# Standard Potentials In Aqueous Solution

**Padhraic Smyth**



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