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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + 2 \text{ C}_2\text{H}_2(\text{g}) + 5 \text{ O}_2(\text{g}) + 2 \text{ C}_2\text{H}_6(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + 2 \text{ H}_2\text{O}(\text{g}) + 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2(\text{g}) + \text{O}_2(\text{g})$
 $\Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$
 $\Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$
 $\Delta H = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$
 $\Delta H = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$
 $\Delta H = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$
 $\Delta H = 285 \text{ kJ/mole}$

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values to determine the heat of reaction for the reaction c_2h_4

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web thus for reaction 9 2 in which 1 mole of h_2o and 1 mole of oh combine to form 2 moles of h_2o $h_55.8$ kj as shown note that the enthalpy change $h_{products} - h_{reactants}$ is positive if heat is absorbed that is if $h_{products}$

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