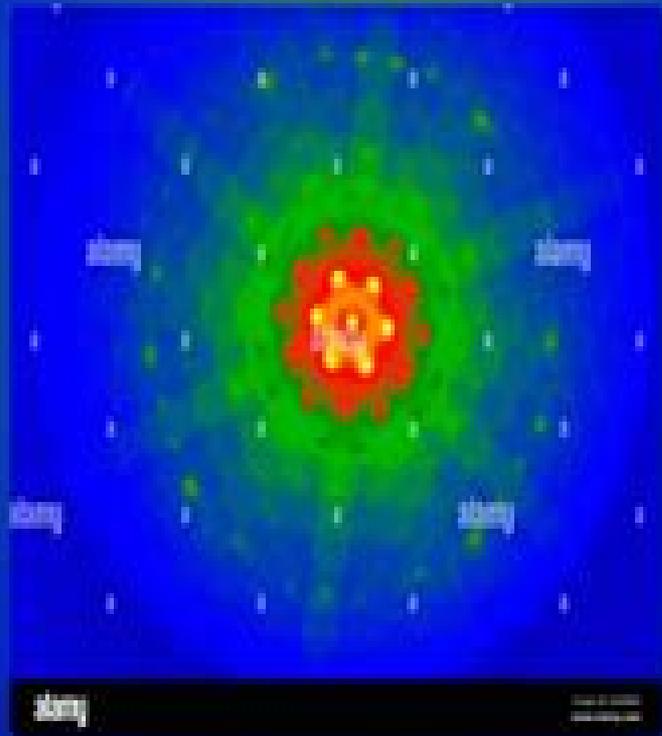


Principles of Electron Diffraction



1 Wave-Particle Duality

Electrons exhibit wave-particle duality, meaning they possess both wave-like and particle-like properties. When electrons interact with a crystal lattice, their wave nature becomes significant.

2 Diffraction

The wave nature of electrons causes them to diffract when they encounter periodic structures, such as the atomic arrangement in a crystal lattice. This diffraction results in a characteristic pattern of interference peaks.

3 Bragg's Law

The positions of these diffraction peaks can be explained by Bragg's law, which relates the wavelength of the electrons, the spacing between atomic planes in the crystal, and the angles of incidence and diffraction.

4 Surface Sensitivity

LEED is sensitive to the outermost atomic layers of a material because the low-energy electrons penetrate only a few atomic layers before being scattered or absorbed.

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