

Valency of First 20 Elements

| Element Name | Symbol | Atomic Number | Electronic Configuration | Lose/gain/share Electrons | Valency |
|--------------|--------|---------------|--------------------------|---------------------------|---------|
| Hydrogen | H | 1 | 1 | Lose or share 1 Electron | 1 |
| Helium | He | 2 | 2 | Stable as per Duplet rule | 0 |
| Lithium | Li | 3 | 2,1 | Lose 1 Electrons | 1 |
| Beryllium | Be | 4 | 2,2 | Lose 2 Electrons | 2 |
| Boron | B | 5 | 2,3 | Lose 3 Electrons | 3 |
| Carbon | C | 6 | 2,4 | Share 4 Electrons | 4 |
| Nitrogen | N | 7 | 2,5 | Gain or share 3 Electrons | 3 |
| Oxygen | O | 8 | 2,6 | Gain or share 2 Electrons | 2 |
| Fluorine | F | 9 | 2,7 | Gain or share 1 Electrons | 1 |
| Neon | Ne | 10 | 2,8 | Stable as per Octet rule | 0 |
| Sodium | Na | 11 | 2,8,1 | Lose 1 Electrons | 1 |
| Magnesium | Mg | 12 | 2,8,2 | Lose 2 Electrons | 2 |
| Aluminum | Al | 13 | 2,8,3 | Lose 3 Electrons | 3 |
| Silicon | Si | 14 | 2,8,4 | Share 4 Electrons | 4 |
| Phosphorous | P | 15 | 2,8,5 | Gain or share 3 Electrons | 3 |
| Sulfur | S | 16 | 2,8,6 | Gain or share 2 Electrons | 2 |
| Chlorine | Cl | 17 | 2,8,7 | Gain or share 1 Electrons | 1 |
| Argon | Ar | 18 | 2,8,8 | Stable as per Octet rule | 0 |
| Potassium | K | 19 | 2,8,8,1 | Lose 1 Electrons | 1 |
| Calcium | Ca | 20 | 2,8,8,2 | Lose 2 Electrons | 2 |

Valency And Molecular Structure

Faraday Society



Valency And Molecular Structure:

Valency and Molecular Structure E. Cartmell, G. W. A. Fowles, 2013-10-22 *Valency and Molecular Structure* Fourth Edition provides a comprehensive historical background and experimental foundations of theories and methods relating to valency and molecular structures In this edition the chapter on Bohr theory has been removed while some sections such as structures of crystalline solids have been expanded Details of structures have also been revised and extended using the best available values for bond lengths and bond angles Recent developments are mostly noted in the chapter on complex compounds while a new chapter has been added to serve as an introduction to the spectroscopy of complex compounds Other topics include the experimental foundation of the quantum theory molecular orbital method ionic hydrogen and metallic bonds structures of some simple inorganic compounds and electronic spectra of transition metal complexes This publication is a useful reference for undergraduate students majoring in chemistry and other affiliated science subjects

Valency and Molecular Structure Edward Cartmell, 1979 *Valency and Molecular Structure* Edward Cartmell, Gerald Wilfred Albert Fowles, 1961 **Journal of the American Chemical Society** American Chemical Society, 1914 Proceedings of the Society are included in v 1 59 1879 1937

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