

Thermodynamics

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Basic Definition of Thermodynamics

The Oxford English Dictionary says: "Thermodynamics: the theory of the relations between heat and mechanical energy, and of the conversion of either into the other." In simpler terms, we can think of thermodynamics as the science that tells us which minerals or mineral assemblages will be stable under different conditions.

In practical terms, thermodynamics not only allows us to predict what minerals will form at different conditions (*forward modeling*), but also allows us to use mineral assemblages and mineral compositions to determine the conditions at which a rock formed (*thermobarometry*). The calculations are often complex and are best carried out using *thermodynamic modeling programs* or programs specifically designed for thermobarometry.

The Basis for Thermodynamic Calculations

All phases, whether mineralogical or not, have an associated **Gibbs Free Energy of Formation** value abbreviated ΔG_f . The ΔG_f value describes the amount of energy that is released or consumed when a phase is created from other phases.

We can calculate the Gibbs Free energy of any reaction (ΔG_{rxn}) by summing the energies of the right-hand side of the reaction and subtracting the energies of the left hand side. If the calculation reveals that $\Delta G_{\text{rxn}} < 0$, the reaction proceeds to the right. If the $\Delta G_{\text{rxn}} > 0$, the reaction proceeds to the left.

Consider the reaction albite = jadeite + quartz (Figure 1). Under normal earth surface conditions, the Gibbs Energy of this reaction is greater than zero. Consequently albite is stable and the assemblage jadeite + quartz is unstable.

The ΔG_f of a mineral varies with changes in pressure (P), temperature (T) and mineral composition (X). Consequently, the ΔG_{rxn} for any reaction will vary with P, T and X.

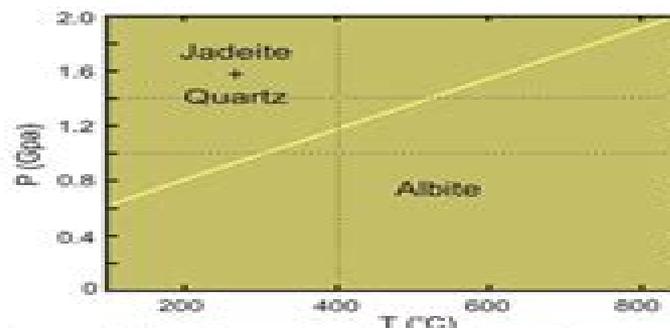


Figure 1. Phase diagram showing the reaction albite = jadeite + quartz. Along the reaction line, the Gibbs Energy of Reaction, $\Delta G_{\text{rxn}} = 0$. Above the line $\Delta G_{\text{rxn}} < 0$; below the line $\Delta G_{\text{rxn}} > 0$.

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