

# Hand- Manipulated Stitches for Machine Knitters



*Susan Guagliumi*

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# The Machine Knitters Handbook

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web as a result of the reaction the equilibrium concentration of  $NH_3(g)$  became  $0.540\text{ M}$  what is the value of the equilibrium constant for this reaction at the given temperature of  $700\text{ C}$   $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$  i write the initial concentrations of reactants and products

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web equilibrium  $1.864\text{ M}$   $1.364\text{ M}$   $0.636\text{ M}$  equilibrium  $1.600\text{ M}$   $1.000\text{ M}$   $0.400\text{ M}$  trial 5 a  $g$   $b$   $g$   $c$   $g$  trial 6 a  $g$   $b$   $g$   $c$   $g$  initial  $1.400\text{ M}$   $1.200\text{ M}$   $0.400\text{ M}$  initial  $0.750\text{ M}$   $2.000\text{ M}$   $0.250\text{ M}$  change change equilibrium  $1.388\text{ M}$   $1.188\text{ M}$   $0.412\text{ M}$  equilibrium  $0.675\text{ M}$   $1.925\text{ M}$   $0.325\text{ M}$  1 examine model 1 a

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