



Sophia Vari

Nohra Haime Gallery (New York)



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 pK_b of the weak acid or base being titrated a solution ph as a function of the volume of 1.00 M NaOH added to 10.00 ml of
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ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $ph = -\log 0.100 = 1.000$ ph = 0.
100 = 1.000 b titrant volume 12.50 ml

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