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AN OBBY**



Roblox Obby Guide

WJ Hussar



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was produced i what mass of iodine was produced

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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g Al $\times 1$ mol Al $\times 2$ mol Al_2O_3 $\times 101.964$ g Al_2O_3 302.3 g Al_2O_3 26.982 g Al 4 mol Al 1 mol Al_2O_3 c determine the percent yield
yield actual $\times 100$ 260.0 g $\times 100$ 86.01 theoretical 302.3 g

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C is the limiting reagent 4.33 g of H 2 are left over 13

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yield theoretical $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ based on this definition we would expect a percent yield

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CH₃OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH₃OH if 68.5 g of CO is
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