

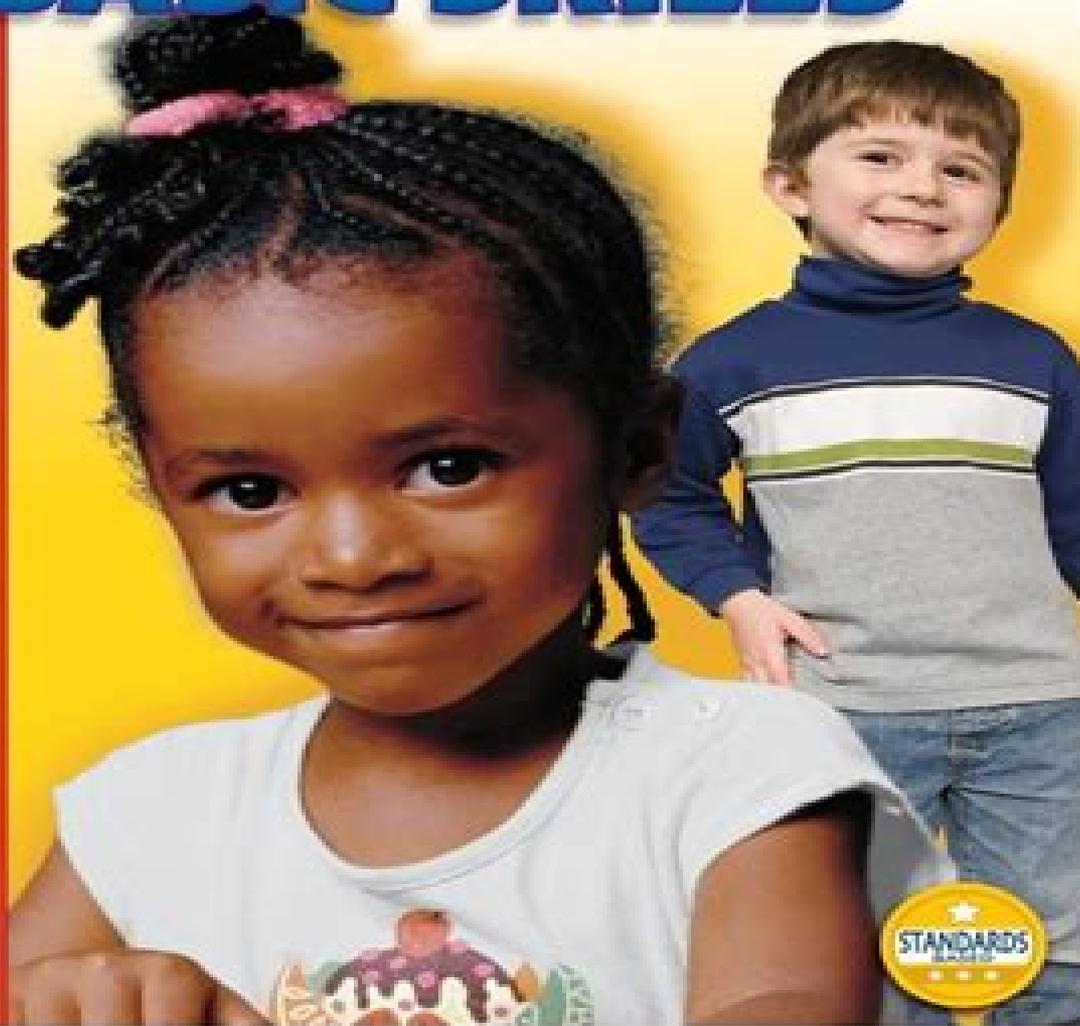
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# Total Basic Skills

**Ying Liu**



**Total Basic Skills:**

## Unveiling the Magic of Words: A Report on "**Total Basic Skills**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Total Basic Skills**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound affect on the souls of its readers.

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{Cu(NO}_3)_2$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H<sub>2</sub>S and multiply it with 2 mole H<sub>2</sub>O over 2 mole H<sub>2</sub>S from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

8.3 mole H<sub>2</sub>S x 18g H<sub>2</sub>O / mole H<sub>2</sub>O = 149.4g H<sub>2</sub>O is your theoretical yield so percent yield = 137.1g H<sub>2</sub>O actual yield / 149.4g H<sub>2</sub>O x 100 = 91.77%

problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O<sub>2</sub> produces a 47.3% yield how many grams of CO would be produced

step one figure out how many moles of the limiting reagent you have

$$21.3 \text{ g O}_2 \times \frac{1 \text{ mole O}_2}{32 \text{ g O}_2} = 0.666 \text{ moles O}_2$$

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$$\text{yield efficiency} = \frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$$

in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction

$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub>

$$160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$$

c determine the percent yield

$$\frac{\text{yield actual}}{\text{yield theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$$

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