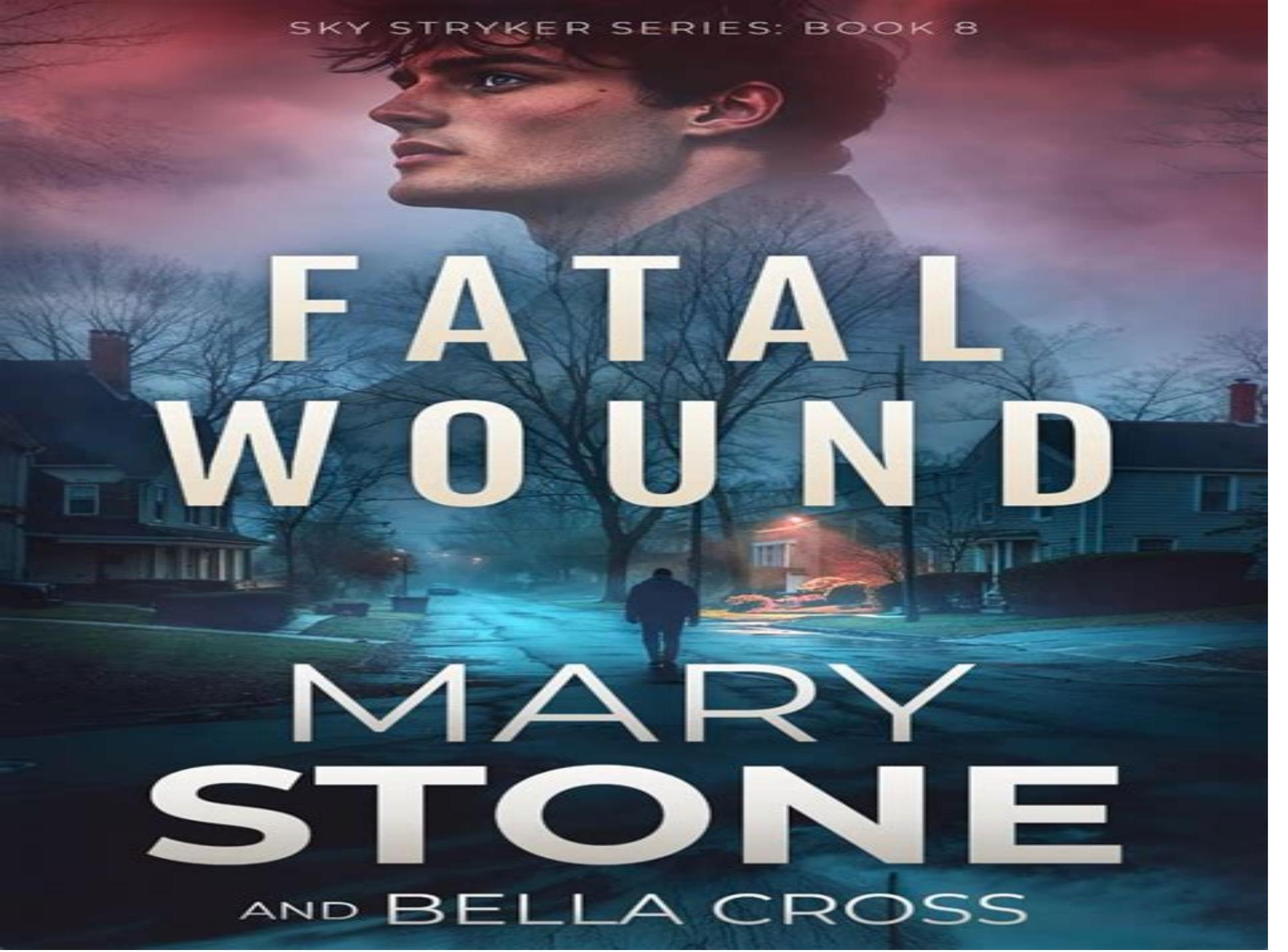


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web the following electron configurations belong to which elements 21 $1s^2 2s^2 2p^6 3s^1$ sodium 22
 $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^6$ ruthenium 23 kr $5s^2 4d^{10}$ cadmium 24 xe $6s^2 4f^{14} 5d^{10} 6p^2$ lead 25 rn $7s^2 5f^{14} 6d^4$
seaborgium determine if the following electron configurations are correct 26 $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^6 5s^1$ no it should
be $3d^{10}$

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each isotope answer co has 27 protons 27 electrons and 33 neutrons $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^7$ i has 53 protons 53
electrons and 78 neutrons $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 5s^2 5p^5$

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web select the electron configuration tab on show number of electrons introduction beyond argon it is a bit tricky to
determine which subshell gets filled next there are several rules that scientists use to determine the electron configurations
of larger atoms question how are the electron configurations of elements beyond argon determined