



kinetics

activation energy

$$\Delta G^\ddagger = G_{TS^\ddagger} - G_{\text{reactants}}$$

$$\text{or } E_a = E_{TS^\ddagger} - E_{\text{reactants}}$$

controls rate of reaction

thermodynamics

$$\Delta G = G_{\text{products}} - G_{\text{reactants}}$$

$$\text{or } \Delta E = E_{\text{products}} - E_{\text{reactants}}$$

$$\Delta H = H_{\text{products}} - H_{\text{reactants}}$$

controls equilibrium K

Techniques In Organic Reaction Kinetics

Michal Rosen-Zvi



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