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web a reaction and its experimental rate law are represented below $2 \text{ ICl}_g + \text{H}_2 \text{ g} \rightarrow 2 \text{ HCl}_g$ rate $k [\text{ICl}]^2 [\text{H}_2]$ a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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a sample of Mg ribbon is added to an erlenmeyer flask containing 50 ml of 3.0 M HCl at 25 °C

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