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The Schumpeterian System

RC Schank



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 $2 \text{ CO}_2 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + \frac{5}{2} \text{O}_2 \text{ (g)}$ $\Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O (g)} + \text{H}_2 \text{ (g)} + \frac{1}{2} \text{O}_2 \text{ (g)} \rightarrow \text{C}_2\text{H}_6 \text{ (g)}$ $\Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (g)}$ $\Delta H = 283 \text{ kJ}$
1 $\text{C}_2\text{H}_6 \text{ (g)}$

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302 g 2 co 2 g 2 h 2 o l h 1411 kj mole c 2 h 6 g 7 2 o 2 g 2 co 2 g 3 h 2 o l h 1560 kj mole h 2 g 1 2 o 2 g h 2 o l h 285 kj mole

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web thus for reaction 9 2 in which 1 mole of H_2O_2 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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