

# THINGS I NEVER LEARNED IN SCHOOL

- How to do taxes
- What taxes are
- How to manage money
- Anything to do with banking
- How to grow food
- How to handle rejection
- How to love myself
- How to handle interviews
- How to buy my first home
- How to survive in the wild
- How to handle myself in dangerous situations

**But I sure am glad I know  
the Pythagorean Theorem**

# Things I Never Learned At School

**Joacim Rocklöv**



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web limiting reactant theoretical and percentage yields key  $4 KO_2 + 2 H_2O \rightarrow 4 KOH + O_2$  g if a reaction vessel contains 0 mol  $KO_2$  and 0 mol  $H_2O$  what is the limiting reactant

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web start with 8.3 mole H<sub>2</sub>S and multiply it with 2 mole H<sub>2</sub>O over 2 mole H<sub>2</sub>S from the equation 8.3 mole H<sub>2</sub>S × 2 mole H<sub>2</sub>O / 2 mole H<sub>2</sub>S × 18g H<sub>2</sub>O = 149.4g H<sub>2</sub>O is your theoretical yield so percent yield = 137.1g H<sub>2</sub>O actual yield / 149.4g H<sub>2</sub>O × 100 = 91.77 problems 6 & 8 are similar to problem 5

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web 1.csun.edu/hcchm001/introchemhandouts/html/limiting\_reagents\_theoretical\_actual\_and\_percent\_yields.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH<sub>3</sub> can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation C<sub>3</sub>H<sub>6</sub> + 3O<sub>2</sub> → 3CO + 3H<sub>2</sub>O if the reaction of 21.3 grams of O<sub>2</sub> produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have 21.6g / 32 g/mol = 0.675 moles

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web percent yield is a measure of the efficiency of the experimental design yield efficiency = mass of product obtained / calculated mass of the product expected × 100 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction Al + O<sub>2</sub> → Al<sub>2</sub>O<sub>3</sub> b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub> 160.0 g Al × 1 mol Al / 27.0 g Al × 2 mol Al<sub>2</sub>O<sub>3</sub> / 3 mol Al × 101.964 g Al<sub>2</sub>O<sub>3</sub> / 1 mol Al<sub>2</sub>O<sub>3</sub> = 269.82 g Al<sub>2</sub>O<sub>3</sub> c determine the percent yield yield actual / yield theoretical × 100 = 260.0 g / 269.82 g × 100 = 96.73%

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web 7Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield = 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H<sub>2</sub> are left over 13

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web the percent yield is calculated as follows text percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield based on this definition we would expect a percent yield

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