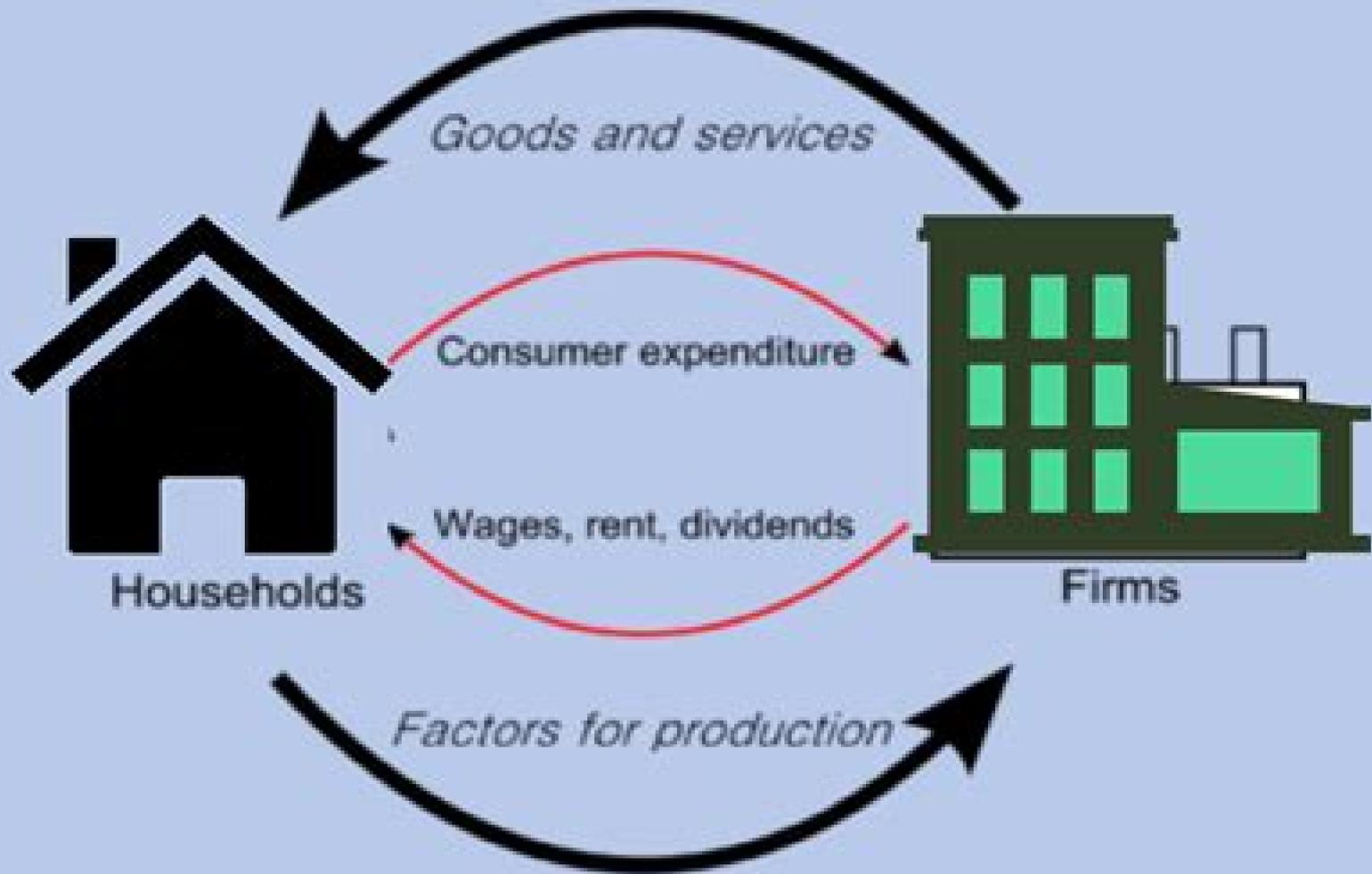


# Two sector Model



# Two Performances detour

**M Lipman**



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web nov 6 2023  $x_a \text{ mol } a \text{ mol } a \text{ mol } b$  and  $x_b \text{ mol } b \text{ mol } a \text{ mol } b$  note in the given mixture the sum of all the mole fractions

present is equal to one  $x_a x_b = 1$  when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is a unitless and dimensionless expression image will be uploaded soon

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just

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web since the mole fraction of the O<sub>2</sub> is 0.050 it contributes 0.050 or if you will 5.0% of the total pressure thus total pressure is 0.210 / 0.050 = 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of O<sub>2</sub> are mixed in a flask at room temperature

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