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Manual Sleep Optimization

Zhe Xu



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quantum number $m_s = \pm \frac{1}{2}$ and $l = 2$

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web the four different types of orbitals s p d and f have different shapes and one orbital can hold a maximum of two electrons the p d and f orbitals have different sublevels thus can hold more electrons as stated the electron configuration of each element is unique to its position on the periodic table

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web explanation the first two electrons fit into the 1s orbital $1s^2$ then next two into the 2s orbital $2s^2$ the 2p orbital has three slots which can take 2 electrons each a total of 6 which gives $2p^6$ the final two electrons fit in the 3s orbital $3s^2$ electron configuration guided notes teacher edition

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certain rules to follow when assigning electrons to different orbitals the periodic table is an incredibly helpful tool in writing electron configurations

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1s22s22p63s23p64s23d104p65s24d6 ruthenium 23 kr 5s24d10 cadmium 24 xe 6s24f145d106p2 lead 25 rn 7s25f146d4 seaborgium determine if the following electron configurations are correct 26 1s22s22p63s23p64s24d104p65s1 no it should be 3d10

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web how many protons neutrons and electrons are in atoms of these isotopes write the complete electron configuration for each isotope answer co has 27 protons 27 electrons and 33 neutrons 1s 2 2s 2 2p 6 3s 2 3p 6 4s 2 3d 7 i has 53 protons 53 electrons and 78 neutrons 1s 2 2s 2 2p 6 3s 2 3p 6 3d 10 4s 2 4p 6 4d 10 5s 2 5p 5

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of larger atoms question how are the electron configurations of elements beyond argon determined