

BUSINESS DECISION-MAKERS GUIDE: RENEWABLE ENERGY IN SOUTH AFRICA



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ZnCl_2 Cu Zn CuCl_2 ZnCl_2 Cu how many moles of Cu ZnCl_2 ZnCl_2 will be produced from 23.0 g of Cu Zn Zn assuming CuCl_2 CuCl_2 is available in excess

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 $\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$
 $2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$
 $3\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2 \rightarrow 4\text{H}_2\text{O} + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O}$
 $5\text{NaBr} + 3\text{SnO}_2 + 3\text{NF}_3 + 3\text{SnF}_2 + 1\text{N}_2\text{O}_3 \rightarrow$ using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $5\text{NaBr} + 3\text{SnO}_2 + 3\text{NF}_3 + 3\text{SnF}_2$

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answer $\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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 $3\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2 \rightarrow 4\text{H}_2\text{O} + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O}$
 $5\text{NaBr} + 3\text{SnO}_2 + 3\text{NF}_3 + 3\text{SnF}_2 + 1\text{N}_2\text{O}_3 \rightarrow$ solve the following stoichiometry grams grams problems 6

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