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**SUPER FUN ACTIVITY
& COLORING BOOK**



Super Fun Activity 1

Jianjun Gao



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web limiting reactant theoretical and percentage yields key $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
 8.3 mole H₂S x 18g H₂O / mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts/html/limiting_reagents_theoretical_actual_and_percent_yields.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N₂ what is the limiting reagent if 8.52 g are actually formed what is the

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 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$ if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have
 $21.3 \text{ g O}_2 \times \frac{1 \text{ mole O}_2}{32 \text{ g O}_2} = 0.6656 \text{ moles O}_2$
 $0.6656 \text{ moles O}_2 \times \frac{3 \text{ moles CO}}{3 \text{ moles O}_2} = 0.6656 \text{ moles CO}$
 $0.6656 \text{ moles CO} \times 28 \text{ g CO} = 18.64 \text{ g CO}$

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web percent yield is a measure of the efficiency of the experimental design yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
 $4 \text{ Al} + 3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$
 b determine the theoretical yield of Al₂O₃
 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.964 \text{ g Al}_2\text{O}_3 = 152.9 \text{ g Al}_2\text{O}_3$
 $302.3 \text{ g Al}_2\text{O}_3 \times \frac{1 \text{ mol Al}_2\text{O}_3}{302.3 \text{ g Al}_2\text{O}_3} \times 26.982 \text{ g Al} = 26.982 \text{ g Al}$
 c determine the percent yield
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$
 theoretical = 302.3 g

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web 7 Na₂C₂O₄ is the limiting reactant percent yield = $\frac{86.68}{100} = 86.68\%$ only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H₂ are left over 13

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 $80\text{ g I}_2\text{O}_5 \times \frac{1\text{ mol I}_2\text{O}_5}{1\text{ mol I}_2\text{O}_5} \times \frac{1\text{ mol I}_2}{1\text{ mol I}_2\text{O}_5} \times \frac{333.8\text{ g I}_2}{1\text{ mol I}_2\text{O}_5} = 28\text{ g CO} \times \frac{1\text{ mol CO}}{1\text{ mol CO}}$

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