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 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} \times 18 \text{ g H}_2\text{O} = 149.4 \text{ g H}_2\text{O}$ is your theoretical yield so percent yield
 $\frac{137.1 \text{ g H}_2\text{O}}{149.4 \text{ g H}_2\text{O}} \times 100 = 91.77\%$ problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts/html/limiting_reagents_theoretical_actual_and_percent_yields.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$ if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have
 $0.216232 \text{ grams mole}^{-1} \times 21.3 \text{ grams} = 0.66 \text{ moles}$ 32 g/mol

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained / calculated mass of the product expected x 100 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 Al + 3 O₂ → 2 Al₂O₃ b determine the theoretical yield of Al₂O₃
 $160.0 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$ 26.982 g/mol Al
c determine the percent yield
 $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$ theoretical 302.3 g

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H₂ are left over 13

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