

Test Taking Strategies

It is normal to feel a little bit nervous or stressed before a test. Almost everyone does! These proven test-taking strategies can help you feel less stressed and do better on your tests.



#1 Do the Word Problem Questions First

Before you look at any of the other questions, find the word problems, and do these questions first. Word problems take a lot of brain power, more time than simple calculations or multiple choice. So it's best to tackle them when your brain is fresh at the beginning of the test.

#2 Take your time

Don't be this guy! Remember to take your time and double-check your answers. Checking your answers means going back and making sure that your solutions make sense. First, check your multiplication and other operations to be sure you read the question correctly. Then, plug your answer back into the equation if you have solved anything algebraically and see if it works.



#3 Manage your time



Be sure you know how much time is left. For example, if you have 60 minutes to answer 30 multiple choice questions, you have about two minutes for each question. So check the clock and make sure you have enough time to complete every question. Just remember - Do not rush through the test!

#4 Read all questions thoroughly

Always read the question thoroughly and carefully. Make sure you know and understand what the question is asking you. If you don't, raise your hand and ask your teacher for assistance.

#5 Eliminate Answers



Most multiple-choice tests have 2 similar answers and 2 answers that do not make any sense. Eliminate the choices that make no sense by marking a line through them. Then you will only have 2 answer choices to choose from. Solve the problem and choose the best one. Remember that you can also use the answers to help you solve the problem if you are stuck and unsure of what to do.

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab

questions write the balanced equation

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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