



# Roblox Tycoon Guide

**J Elliott**



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temperature change plugging in the values given in the problem you get  $q_{\text{water}} = 4.18 \text{ J/g} \times 110 \text{ g} \times 26.6 \text{ C} - 25.0 \text{ C} \times q_{\text{water}}$   
 $550 \text{ J} = \Delta h \times q_{\text{water}} = 550 \text{ J}$

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web calorimetry problems calorimetry problems  $q_{\text{sur}} = m \times c \times \Delta t$  heat  $m$  mass  $q_{\text{rxn}}$   $q_{\text{sur}}$   $t_f$   $t_i$  name per date  $c$  specific heat for water  $4.184 \text{ J/g} \times 1$  what is the specific heat of aluminum if the temperature of a  $28.4 \text{ g}$  sample of aluminum is increased by  $8.1 \text{ C}$  when  $207 \text{ J}$  of heat is added  $2$

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 $14.73 \text{ J} = \Delta h_{\text{rxn}} = q_{\text{rxn}} = \text{calorimeter} = m \times c \times \Delta t$

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web use the information in this chart to answer the following questions 1 if i have  $125 \text{ grams}$  of water at a temperature of  $25.0 \text{ C}$  and heat it to  $145.0 \text{ C}$  what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the process in 1 3 i have  $18.7 \text{ grams}$  of an unknown liquid chemical at a temperature of  $35.0 \text{ C}$

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web sample problem calorimetry and enthalpy changes in an experiment  $25.0 \text{ ml}$  of  $1.00 \text{ M HCl}$  at  $25.0 \text{ C}$  is added to  $25.0 \text{ ml}$  of  $1.00 \text{ M NaOH}$  at  $25.0 \text{ C}$  in a foam cup calorimeter a reaction occurs and the temperature rises to  $32.0 \text{ C}$  calculate the enthalpy change in  $\text{kJ}$  for this reaction assume the densities of the solutions are  $1.00 \text{ g/ml}$  and that

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web before we practice calorimetry problems involving chemical reactions consider a simple example that illustrates the core idea behind calorimetry suppose we initially have a high temperature substance such as a hot piece of metal  $m$  and a low temperature substance such as cool water  $w$  answer the initial temperature of the copper

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web choose 1 answer the final temperature is closer to  $T_1$  than to  $T_2$  a the final temperature is closer to  $T_1$  than to  $T_2$  the final temperature is exactly halfway between  $T_1$  and  $T_2$  b the final temperature is exactly halfway between  $T_1$  and  $T_2$  the final temperature is closer to  $T_2$  than to  $T_1$  c

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