

Algebra Study Sheet

Basic Properties and Identities

Arithmetic and Fraction Operations

$$\begin{aligned}\frac{a}{b} + \frac{c}{b} &= \frac{a+c}{b} \\ \frac{a}{b} - \frac{c}{b} &= \frac{a-c}{b} \\ \frac{a}{b} \cdot \frac{c}{d} &= \frac{ac}{bd} \\ \frac{a}{b} \div \frac{c}{d} &= \frac{a}{b} \cdot \frac{d}{c} = \frac{ad}{bc} \\ \frac{a+b}{c} &\neq \frac{a}{c} + \frac{b}{c} \text{ unless simplified}\end{aligned}$$

Exponent Rules

$$\begin{aligned}a^m \cdot a^n &= a^{m+n} \\ \frac{a^m}{a^n} &= a^{m-n}, \quad a \neq 0 \\ (a^m)^n &= a^{mn} \\ (ab)^n &= a^n b^n \\ \left(\frac{a}{b}\right)^n &= \frac{a^n}{b^n}, \quad b \neq 0 \\ a^0 &= 1, \quad a \neq 0 \\ a^{-n} &= \frac{1}{a^n}\end{aligned}$$

Radical Properties

$$\begin{aligned}\sqrt[n]{ab} &= \sqrt[n]{a} \cdot \sqrt[n]{b} \\ \sqrt[n]{\frac{a}{b}} &= \frac{\sqrt[n]{a}}{\sqrt[n]{b}}, \quad b \neq 0 \\ \sqrt{a^2} &= |a|\end{aligned}$$

Study Guide Beginning Algebra Third Edition

Daniel F McAuley



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g})$ $\Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g})$ $\Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) \rightarrow \frac{7}{2} \text{O}_2(\text{g}) + 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$ $\Delta H = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$ $\Delta H = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$ $\Delta H = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta H = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note

that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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